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Atoms are extremely small. They can be small as 10-8 centimeters or 0.1 naked, in diameter. Atoms are made of three different types of subatomic particles: neutrons, protons and electrons. The core, or the center of an atom, is made of protons, which are particles and positively charged neutrons, which are neutral (have no cost). Electrons are particles loaded negatively. Each atom has the same amount of protons and electrons, so each atom has a neutral charge. During the number of neutrons, protons and atoms of electrons, their properties vary. Which element is an atom is determined by how many protons has. The number of protons in a given atomic core is called its atomic number, and is these atomic numbers with which the elements on the periodic table are classified. For example, a hydrogen atom has a proton in its nucleus and is therefore called number one on the periodic table. Helium has two protons, so his atomic number is two, and so on.Isotopes are atoms that seem and act in the same way but have a different number of neutrons in their nuclei. For example, the atomic hydrogen number is always one because all hydrogen atoms have a proton. However, various hydrogen isotopes, such as H-2, have a proton and neutron; H-3 has a proton and two neutrons, etc. The sum of protons and neutrons in the core of an atom is its atomic mass. Therefore, the atomic mass of iSotope H-2 is two, the atomic mass of IsoTope H-3 is three, and so on. Correct: 1 How to estimate What amount of payment at home fits your budget 2 How much does a gram on a ladder? 3 Here's how you can keep your pets safe during the Covid-19 4 pandemic because it is vaccinated does not mean that you should launch the mask - still 5 that will benefit from the American Jobs infrastructure plan of Biden? 1 How the complicated history of Hong Kong Riverbera today 2 where you can find the HM Surf in Pokemon Platinum? 3 What are the most common bonus ball numbers? 4 How the Covid-19 pandemic has changed the workplaces in lasting ways 5 What is a 5G network and how is 5G from 4G? 1 What did you call a group of sheep? 2 Which country does the country code of 35? 3 What is the Visa H-1B program? 4 How am I tied to my nephew's son? 5 Oil will be kicked in the sidewalk after the collapse of Covid-19 sales? 1 How to estimate What amount of home payment fits your budget 2 How much does a gram on a ladder? 3 Here's how you can keep your pets safe during the Covid-19 4 pandemic because it is vaccinated does not mean that you should launch the mask - still 5 that will benefit from the American Jobs infrastructure plan of Biden? Albedo-ukr / wikimedia commons / cc by-sa 2.5 neon is the element that is atomic number 10 on the periodic table. Lithium is the element that is the atomic number 3 on the periodic table. These mean that every atom contains 3 protons. Lithium is a soft, silvery and light alkaline metal, denoted with the symbol there. Here are interesting facts on atomic number 3: Lithium is the lighter metal and the lighter solid element at temperature and routine pressure. The density of the maximum nearby room temperature is 0.534 g / cm3. This means that not only floats on the water, but it's just a thick half like. It is so light, it can also float on the oil. It also has the maximum specific thermal capacity of a solid element. The number 3 element has the highest fusion point and the boiling point of alkaline metals. The number 3 element is soft enough to cut with shears. Freshly cut metal is silver colored with a metallic shine. However, the wet air quickly corrodes the metal, pinwheel matte gray and finally black. Among his uses, lithium is used in drugs for bipolar disorder, to make them Lithium ions and to add a red color to fireworks. It is also used in glass and ceramics and to grease high temperature lubricant. It is a refrigerant in breeders reactors and a tritium source when the atomic number 3 is bombed by neutrons. Lithium is the only alkaline metal that reacts reacts nitrogen. And yet, it is the least reactive metal in its group of elements. This because the Lithium Valence electron is so close to the atomic nucleus. While lithium metal burns in the water, it doesn't vigorously do as sodium or potassium. The lithium metal will burn in the air and should be preserved under the kerosene or in an inert atmosphere, like Argon. Don't try to turn off a lithium fire with water because you will only worse! Because the human body contains a lot of water, the lithium will also burn the skin. A corrosive and should not be managed without protective gears. The name for the element comes from the Greek word "Lithos", which means "stone". Lithium was discovered in the mineral packet (Liisi4010). Naturalist and Brazilian statesman, JozÁ© BonifÁ©cio de Andrada and Silva found the stone on the Swedish island. Although the mineral looked like a normal gray rock, he was inflamed in red when he thrown into a fire. Swedish chemist Johan August Arfvedon determined that the mineral contained a previously unknown element. He could not isolate a pure specimen, but produced a lithium salt from Petalandite in 1817. The atomic mass of the lithium is 6.941 . The atomic mass is a weighted average that represents the natural abundance of the element's isotope. It is believed that the lithium is one of the only three chemical elements produced in the Big Bang that formed the universe. The other two elements are hydrogen and helium. However, lithium is relatively unusual in the universe. Scientists believe that reason is that lithium is almost unstable, with isotopes that have the lowest bonding energies per nucleus of any stable nuclides. Several lithium isotopes are known, but the natural element is a mix of two stable isotopes. Li-7 (92.41% natural abundance) and Li-6 (natural abundance of 7.59 percent). The most stable radioisotope is Lithium-8, which has a half-life of 838 ms. The lithium easily loses its external electron to form the ion Li +. This leaves the atom with a stable internal shell of two electrons. Lithium ion easily leads electricity. Due to its high reactivity, lithium is not in nature as a pure element, but the ion is abundant in sea water. Lithium compounds are found in clay. The first merger reaction of humanity involved the number 3 of the number 3, in which the lithium was used to make hydrogen isotopes for the merger of Mark Oliphant in 1932. Lithium is found in quantities of traces in the living organisms , but its function is not clear. Lithium salts are used to treat bipolar disorder, where they act to stabilize mood. The lithium is an ordinary pressure superconductor at an extremely low temperature. Also, supercondetti at higher temperatures when the pressure is very high (greater than 20 GPa). Lithium displays multiple structures and crystal allotropics. It exhibits a Rhombohedral crystal structure (nine layered repetition escapement) around 4 k (Elio liquid temperature), passing to a cubic cubic structure and fitted fitted on the body while the temperature increases. Boron is the element that is the atomic number 5 on the periodic table. It is a metal or semimetal that is a solid glossy black at room temperature and pressure. Here are some interesting facts about Boron. Atomic number: 5 Element name: Boron Element Symbol: B Atomic Weight: 10.81 Category: Metallaid Group: Group 13 (Boron Group) Period: Period 2 The boron compounds constitute the base for the classic melma recipe, which polymerizes the compound borax . The name of the Boron element comes from the Arab Word Buraq, which means white. The word was used to describe borax, one of the boron compounds known to ancient man. An Atom Boron has 5 protons and 5 electrons. Its average atomic mass is 10.81. The natural boron consists of a mix of two isotopes Boron-10 and Boron-11. Eleven isotopes, with masses of 7 to 17 are known. Boron presents metal or non-metal properties, depending on the conditions. The number 5 element is present in the cellular walls of all plants, so the plants, as well as any animal eating eating Contain boron. Boron Elemental is not toxic to mammals. Over a hundred minerals contain boron and is found in different compounds, including boric acid, borace, boratori, Kernite and Ulexite. However, pure boron is extremely difficult to produce and the abundance of the element is only 0.001% of the earth's crust. The atomic element number 5 is rare in the solar system. In 1808, Boron was partially purified by Sir Humphry Davy and also by Joseph L. Gay-Lussac and L. J. ThÁ©© Ard. They reached the purity of about 60%. In 1909 Ezekiel Weintraub isolated almost the pure element number 5. Boron has the highest fusion point and the boiling point of metals. The crystalline boron is the second most difficult element, following carbon. Boron is resistant and heat resistant. While many elements are produced by nuclear fusion within stars, Boron is not among them. The boor seems to have been formed by nuclear fusion from collisions of cosmic rays, before the solar system is formed. The amorphous phase of Boron is reactive, while the crystalline boron is not reactive. There is a boron-based antibiotic. It is a derivative of streptomycin and is called Boromycin. Boron is used in surfaces surfaces, magnets, nuclear reactor shielding, semiconductors, to make borosilicate, ceramic, insecticide, disinfectant, detergent, cosmetics and many other products. Boron is added to steel and other alloys. Because it is an excellent neutron absorber, it is used in the nuclear reactor control bars. Atomic number 5 burns with a green flame. It can be used to produce green fire and is added as a common dye in fireworks. Boron can transmit part of infrared light. Boron forms stable covalent bonds rather than ionic obligations. At room temperature, Boron is a bad electrical conductor. Its conductivity improves how it is heated. Although boron nitride is not so difficult as the diamond, it is preferred for use in high temperature equipment because it has a higher thermal and chemical resistance. Boro nitride also form nanotubes, similar to those formed by carbon. However, unlike carbon nanotubes, boron nitride tubes are electric insulators. Boron was identified on the surface of the Moon and Mars. The detection of both waters that Boron on Mars supports the possibility that Mars may have been habitable, at least in the crater gale, at some point in the distant past. The average cost of the pure crystalline boron was around \$ 5 per gram in 2008. Dunitz, J. d.; Hawley, D. M.; Miklos, d.; White, D. N. J.; Berlin, Y.; Marusa "Á© ¤, R.; Prelog, V. (1971). "Boromycin structure". Acta chemical Helvetica. 54 (6): 1709 - 1713. doi: 10.1002 / hlca.19710540624 Eremets, M. I.; Struzhkin, V. v.; Mao, h.; HEMBLEY, R. J. (2001). "SuperconduittivitÁ© in Boron". 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